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Note: Questions with (*) mark are from syllabus of class XI. Q.1. A solution containing 2.675 g of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ (molar mass = 267.5 g mol⁻¹) is passed through a cation exchanger. The chloride ions obtained in solution were treated with excess of AgNO_3 to give 4.78 g of AgCl (molar mass = 143.5 g mol⁻¹). The formula of the complex is (Atomic mass of Ag = 108 u)

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