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will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons

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Section 1 Light and Quantized Energy



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Chapter 5 Section 2 Quantum Theory  
and the Atom Chapter 5 Section 3  
Electron Configuration Filling orbitals

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Chapter 5 - Electrons in Atoms Section  
5.1 - Models of the Atom The  
Rutherford's model of the atom did not

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explain how an atom can emit light or the chemical properties of an atom.

Plum Pudding Model Rutherford's Model

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Revising the Atomic Model 5.2 Electron  
Arrangement in Atoms 5.3 Atomic  
Emission Spectra and ... CHEMISTRY &  
YOU Explain why the correct electron  
configuration of oxygen is  $1s^2 2s^2 2p^4$

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and not  $1s^2 2s^2 2p^3 3s^1$ . The 2p orbitals are lower in energy

### **5.2 Electron Arrangement in Atoms**

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Question 1. What will be the minimum

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pressure required to compress 500 dm<sup>3</sup> of air at 1 bar to 200 dm<sup>3</sup> at 30° C ?

Answer: From the given data :  $P_1 = 1$  bar  
 $P_2 = ?$

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electrons in atoms ...

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Chapter 5 - Electrons in Atoms - 5

Assessment - Page 156: 106 Answer The atomic mass of chlorine is very far from a whole because a weighted average of atomic masses of all of its isotopes is computed in determining its atomic

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mass.

### **Chemistry (12th Edition) Chapter 5 - Electrons in Atoms ...**

In addition to the expected ions Tl  $3+$ , Sn  $4+$ , Pb  $4+$ , and Bi  $5+$ , a partial loss of these atoms' valence shell electrons can also lead to the formation of Tl  $+$ , Sn  $2+$ , Pb  $2+$ , and Bi  $3+$  ions. The formation of

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these 1+, 2+, and 3+ cations is ascribed to the inert pair effect, which reflects the relatively low energy of the valence s -electron ...

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Electrons in Atoms Electron Configurations for Elements in Period Three Table 5-4 Figure 5-19. This sublevel diagram shows the order in which the orbitals are usually filled. The proper sequence for the first seven orbitals is 1s, 2s, 2p, 3s, 3p, 4s, and 3d.

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### **Section Review Answer Key**

Chapter 5 Electrons in Atoms 43 SECTION  
5.1 MODELS OF THE ATOM (pages  
127–132) This section summarizes the  
development of atomic theory. It also  
explains the significance of quantized  
energies of electrons as they relate to  
the quantum mechanical model of the  
atom. The Development of Atomic

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Models (pages 127-128) 1.

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This can be explained because the energy of the subshells increases as  $l$  increases, due to penetration and shielding (as discussed previously in this chapter). Within any one shell, the  $s$

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s electrons are lower in energy than the p electrons. This means that an s electron is harder to remove from an atom than a p electron in the same shell.

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close contacts listings This is just one of the solutions for you to be successful Chapter 11 Modern Atomic Theory 35 Valence electrons are those in the outermost (highest) principal energy level of an

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Gravity. Created by. noahkotlek. Key Concepts: Terms in this set (39) axial position. location in a trigonal bipyramidal geometry in which there is another atom at a  $180^\circ$  angle and the equatorial positions are at a  $90^\circ$  angle ...

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