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12. ChemQuest 15 - Yav Science

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11. In your own words, what is the difference between a sublevel and an orbital? Orbitals are contained within sublevels. A sublevel is a grouping of orbitals. 12. How many electrons can fit in each of the following energy levels: 1st energy level = 2 because only 2 can fit in an s sublevel and the first energy level only has an s sublevel.

Information: Energy Levels and Sublevels

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143 and 144: 143 9. Consider your

answers to que. Page 145 and 146: 145

Critical Thinking Questions 5. Page 147

and 148: 147 8. If a substance is a

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Bronsted. Page 149 and 150: 149 6. Use the information given an. Page 151 and 152: 151 ChemQuest 50 Name: _____
Page 153 and 154: 153 11. Find the pH of a 0.85 M sol

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Answers

The 11th electron goes into the 3s sublevel. The atom with 11 electrons is sodium. 2. a) How many electrons would an atom need to have before it can begin filling the 3d sublevel? 20 electrons because all of these orbitals need to be filled: 1s 2s 2p 3s 3p 4s b) What is the first element that has enough electrons to begin placing ...

Information: Hund's Rule - S.W.H.S CHEMISTRY

Use the information given and your answers to questions 1 and 2 to find the following: a) Divide the pH by the pOH of the solution. $2.60 \div 11.40 = 0.228$ b) Multiply the pH by the pOH of the solution. $2.60 \times 11.40 = 29.64$ c) Add the pH and the pOH together. $2.60 + 11.40 = 14.00$ d) Multiply $[H^+]$ and $[OH^-]$ together. $2.5 \times 10^{-3} \times 4.0 \times 10^{-12} = 1.0 \times 10^{-14}$ e) Divide $[H^+]$ by $[OH^-]$. $2.5 \times 10^{-3} \div 4.0 \times 10^{-12} = 6.25 \times 10^8$

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